



Effective absorption cross sections and photolysis rates of anthropogenic and biogenic secondary organic aerosols



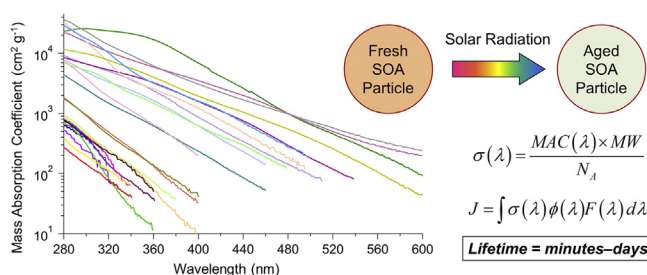
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HIGHLIGHTS

- We measured mass absorption coefficients for a number of different types of SOA.
- We estimated the rates of condensed-phase photochemical processes in SOA.
- These processes can potentially be as fast as oxidation of SOA by gas-phase radicals.

GRAPHICAL ABSTRACT



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ABSTRACT

Mass absorption coefficient (MAC) values were measured for secondary organic aerosol (SOA) samples produced by flow tube ozonolysis and smog chamber photooxidation of a wide range of volatile organic compounds (VOC), specifically: α -pinene, β -pinene, β -myrcene, d-limonene, farnesene, guaiacol, imidazole, isoprene, linalool, ocimene, p-xylene, 1-methylpyrrole, and 2-methylpyrrole. Both low- NO_x and high- NO_x conditions were employed during the chamber photooxidation experiments. MAC values were converted into effective molecular absorption cross sections assuming an average molecular weight of 300 g/mol for SOA compounds. The upper limits for the effective photolysis rates of SOA compounds were calculated by assuming unity photolysis quantum yields and convoluting the absorption cross sections with a time-dependent solar spectral flux. A more realistic estimate for the photolysis rates relying on the quantum yield of acetone was also obtained. The results show that condensed-phase photolysis of SOA compounds can potentially occur with effective lifetimes ranging from minutes to days, suggesting that photolysis is an efficient and largely overlooked mechanism of SOA aging.

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1. Introduction

Primary organic aerosol (POA) directly emitted from anthropogenic and biogenic sources and secondary organic aerosol (SOA) produced by gas-phase or aqueous-phase oxidation of atmospheric

volatile organic compounds (VOC) have both health and climate relevance (Ervens et al., 2011; Finlayson-Pitts and Pitts, 2000; Solomon et al., 2007). Biogenic VOC that contribute the most to SOA production include isoprene, α -pinene, β -pinene, myrcene, and limonene (Geron et al., 2000). In urban areas, major sources of VOC are from fossil fuel combustion as well as industrial emissions (World Health, 1989). Once in the atmosphere, organic aerosol (OA) is known to undergo both physical and chemical aging processes from reactions with sunlight and atmospheric oxidants (George

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et al., 2015). It is important to understand the nature of these changes in order to predict the health and climate impacts of atmospheric aerosols that travel long distances from the point of their initial formation.

Prevalent chemical mechanisms for OA aging involve multi-phase oxidation of OA constituents by atmospheric oxidants. The most dominant oxidant involved in the multiphase oxidation of OA, as well as aqueous-phase oxidation of water-soluble atmospheric organics, is the OH radical (Ervens et al., 2011; George et al., 2015). NO₃, O₃, and other oxidants also contribute to these aging processes. OA compounds can additionally be aged by condensed-phase photochemical reactions occurring in the bulk of the SOA material upon exposure to solar actinic radiation (George et al., 2015). Currently, experimental evidence to determine the relative importance of condensed-phase photochemistry occurring in OA relative to multi-phase oxidation processes is limited. Chamber studies have begun to examine this effect by studying the properties and composition of SOA before and after exposure to UV light. Presto et al. (2005) found that the formation of α -pinene ozonolysis SOA in a smog chamber was suppressed by 20–40% when exposed to UV light. Kroll et al. (2006) observed that low-NO_x isoprene SOA grew during the initial irradiation period, but that was followed by a decrease in particle size as the SOA mixture was further irradiated. However, β -pinene low-NO_x SOA particles did not shrink upon irradiation, indicating that the photochemical effects of UV exposure are dependent on the chemical composition of the SOA. Surratt et al. (2006) detected a similar decrease in the particle size of low-NO_x isoprene SOA upon irradiation. In these experiments, it was difficult to separate the effects of gas-phase and condensed-phase photochemistry. To bypass this limitation, Epstein et al. (2014) used a denuder to remove VOC from α -pinene ozonolysis SOA before the irradiation. They observed that UV light slightly reduced both the particle mass concentration and the average particle size; the fraction of organic peroxides in the particles decreased upon irradiation. Epstein et al. (2014) and Wong et al. (2014) similarly observed a reduction in the particle size for α -pinene ozonolysis SOA material that was collected on a filter and re-aerosolized back in a photolysis chamber; the effect was larger at elevated relative humidity (RH). In contrast, Kourtchev et al. (2015) observed no significant change to molecular level composition of SOA prepared by ozonolysis of a mixture of four biogenic VOC (α -pinene, β -pinene, Δ 3-carene and isoprene) and exposed to actinic UV radiation. In summary, chamber experiments suggest that photochemical processes in SOA particles are possible, but the results are inconsistent.

Qualitative experiments with irradiation of bulk SOA samples were carried out by Walser et al. (2007) and Mang et al. (2008). Both studies showed that SOA particles formed from limonene ozonolysis absorbed radiation in the actinic region, most likely due to the carbonyl and peroxide functional groups present, and that the ensuing photochemistry generated a number of small VOC as detectable products. The carbonyl functional groups appeared to play a significant role in this aging, as evidenced by the production of measureable quantities of CO, CH₄, acetaldehyde, acetone, and other VOC, many of which could be explained by standard Norrish type-I and Norrish type-II carbonyl photochemistry (Mang et al., 2008; Pan et al., 2009).

Condensed-phase photochemistry is not limited to the OA particles; it can also take place in cloud/fog droplets. The irradiation of aqueous extracts of limonene ozonolysis SOA resulted in a significant change in the molecular composition, in which oligomeric and carbonyl compounds were depleted (Bateman et al., 2011). The photolytic processing appeared to be more efficient when SOA was dissolved in an aqueous solution as compared to dry SOA particles. The direct aqueous photolysis of high-NO_x isoprene SOA also

produced major changes in the molecular composition (Nguyen et al., 2012). On average, 30% of the SOA by mass was significantly modified, suggesting that the effects of direct aqueous photolysis should not be ignored. Aqueous solutions of light-absorbing OA compounds, such as SOA formed by photooxidation of naphthalene under high-NO_x conditions, SOA formed by limonene ozonolysis and aged with ammonia vapor, and products of aqueous reaction of glyoxal and methylglyoxal with ammonium sulfate were shown to efficiently photobleach and undergo a large change in composition upon exposure to irradiation (Lee et al., 2014; Sareen et al., 2010; Zhao et al., 2015). Finally, a systematic study of photochemistry of different types of SOA by Romonosky et al. (2015) showed that aqueous photolysis may be an effective aging mechanism for SOA derived from biogenic, anthropogenic, and biomass burning VOC precursors.

The experiments done so far have been qualitative in nature; the diversity in structure and reactivity of VOC complicate the scientific understanding of the photochemistry of atmospheric aerosols. To better estimate the time scales of these photochemical processes under realistic atmospheric conditions, this study focuses on the quantitative measurements of absorption coefficients of SOA formed through the oxidation of biogenic and anthropogenic VOC using common atmospheric oxidants, specifically OH (in presence or absence of NO_x) and O₃. From these data, the rates of photolysis for a variety of SOA were calculated, resulting in a better understanding of the lifetimes of SOA in the atmosphere with respect to condensed-phase photochemistry.

2. Materials and methods

Both O₃-initiated and OH-initiated oxidation conditions, with and without added NO_x, were used to prepare model SOA from VOC as shown in Table 1 and Table 2. VOC were used at their stated commercial purity (Table S1). All experiments were performed under dry conditions (relative humidity < 2%) and in the absence of seed particles. The O₃-initiated reactions were carried out in a 17 L flow tube reactor described by Bones et al. (2010). Liquid VOC was injected by a syringe pump at a rate of 25 μ L/h into a 5 SLM (standard liters per minute) flow of zero air. A 0.5 SLM flow of oxygen (Airgas; 99.994% purity) was sent through an ozone generator and a custom-made ozone photometric detector. The O₃- and VOC-containing flows were mixed at the entrance of the flow cell. Typical mixing ratios in the flow cell ranged from 60 to 100 ppm (parts per million by volume) of O₃ and 10–30 ppm of VOC, respectively. Residence time in the flow tube was less than 5 min, but sufficient to oxidize most of the VOC and to form SOA. A 1-m charcoal denuder scrubbed the residual O₃ and gas-phase organic compounds from the flow exiting the reactor. Multiple SOA collection methods were used in the experiments. In some experiments, SOA was collected by impaction on foil substrates using a multi-orifice uniform-deposit impactor (MOUDI, MSP model 110-R) sampling at 30 SLM. In other experiments, SOA was collected on stage D of a Sioutas impactor (Sioutas, 2004) modified to accept either a CaF₂ window or foil substrates. The specified cut-off point of this impactor stage was 0.25 μ m at 9 SLM flow rate, but sampling was done at 15 SLM to drive the cut-off point to lower particle sizes. The foil substrates and CaF₂ windows were weighed before and after SOA collection to determine the mass of SOA collected. In several experiments, SOA was collected on polytetrafluoroethylene (PTFE) filters (Millipore fluoropore, 0.2 μ m pore size). Absorption coefficients from SOA collected on foil substrates, CaF₂ windows, and PTFE filters were comparable in magnitude. Depending on the yield of SOA, collection time for the flow tube was approximately 30 min to 2 h. The samples were either analyzed immediately or sealed and frozen for later analysis.

Table 1
Names and abbreviations for the VOC used to generate SOA samples in a smog chamber. The SOA samples are referred to using the code VOC/oxidation conditions; specifically, VOC/OH and VOC/OH/NO_x refer to samples prepared under low-NO_x and high-NO_x conditions, respectively. The reaction time in the chamber is equivalent to the irradiation time.

Precursor (abbreviated name)	Oxidant	Initial VOC (ppm)	Initial NO (ppb)	Reaction time (h)	Collection time (h)	Amount typically collected (mg)
2-Methylpyrrole (2MPYR)	OH	0.5	<1	3	2	1.3
2MPYR	OH/NO _x	0.5	300	2	3	1.8
α -Pinene (APIN)	OH/NO _x	0.5	300	2	3	2.3
Farnesene (FAR)	OH	0.5	<1	1.5	3	3.5
FAR	OH/NO _x	0.5	300	1.5	3	3.0
Guaiacol (GUA)	OH	0.5	<1	2	3	0.95
GUA	OH/NO _x	0.5	300	1	3	1.3
Isoprene (ISO)	OH	4.0	<1	4	2.5	0.10
Linalool (LIN)	OH	0.5	<1	2	3	1.5
LIN	OH/NO _x	0.5	300	2	3	1.5
Naphthalene (NAP)	OH	0.4	<1	2	3	1.0
NAP	OH/NO _x	0.4	400	2	3	1.0
Ocimene (OCI)	OH	0.5	<1	2	3	2.0
OCI	OH/NO _x	0.5	300	2	3	1.2
Toluene (TOL)	OH	1.0	<1	3	3	1.3
TOL	OH/NO _x	1.0	1000	3.5	4	0.70
p-Xylene (XYL)	OH	1.0	<1	3	3	0.60
XYL	OH/NO _x	1.0	300	3	3	1.0

Table 2

Names and abbreviations for the VOC used to generate flow tube ozonolysis SOA samples. The initial VOC concentration in flow tube experiments is the steady-state mixing ratio the VOC would have in the absence of ozone. Ozone was added in small excess with respect to the VOC. The flow tube residence time is on the order of minutes.

Precursor	Initial VOC (ppm)	Amount typically collected (mg)
2MPYR	20	0.80
APIN	10	2.0
ISO	10	0.15
1-Methylpyrrole (1MPYR)	20	0.20
Imidazole (IMID)	30	0.20
β -Myrcene (BMYR)	10	0.40
β -Pinene (BPIN)	10	0.70
d-Limonene (LIM)	10	7.3

The OH-initiated reactions were carried out in a ~5 m³ Teflon chamber, equipped with UV-B lights (FS40T12/UVB, Solarc Systems Inc.) with emission centered at 310 nm. The reactions took place under dry conditions (relative humidity <2%), in the absence of seed particles. Chamber experiments were carried out either under “high-NO_x” oxidation conditions, with approximately 300 ppb (parts per billion by volume) of NO added to the chamber to simulate an urban atmospheric environment, or “low-NO_x” oxidation conditions, with no additional NO added to the chamber in order to simulate a remote atmospheric environment. For the OH precursor, 90 μ L of H₂O₂ (Aldrich; 30% by volume) was injected in the chamber by evaporation under a stream of zero air resulting in ~4 ppm of H₂O₂ vapor in the chamber. Precursor VOC was added to the chamber using the same method as was used for the H₂O₂ addition, with VOC concentrations in the chamber varying from 500 ppb to 1 ppm. One exception was isoprene; we had to increase its starting mixing ratio to 4 ppm in order to generate enough SOA for further analysis. The chamber content was mixed for several minutes using a fan, which was then turned off to minimize particle wall losses. The UV-B lamps were turned on for 1.5–4 h to initiate photochemistry (see Table 1). SOA particle concentration within the chamber was monitored by a TSI model 3936 scanning mobility particle sizer (SMPS). A Thermo Scientific model 49i ozone analyzer recorded O₃ data, while a Thermo Scientific model 42i-Y NO_y analyzer recorded NO/NO_y data. SOA particles passed through an activated carbon denuder at approximately 15 SLM and were collected using PTFE filters. Collection times ranged from 2 to 3 h, with approximately 0.1–3.5 mg of SOA collected on the filter, depending on the sample. The amount of SOA on the filter was estimated from SMPS data,

assuming an SOA particle density of 1.2 g/cm³ and 100% collection efficiency by the filters. This assumed value of particle density represents an average of densities of anthropogenic and biogenic SOA materials reported by Hallquist et al. (2009).

All of the samples were dissolved by vigorous shaking of the substrates immersed in 5 mL of water or methanol solvent for 10–15 min. Sonication was avoided to prevent formation of peroxides in the extract (Mutzel et al., 2013). Water was used as the extracting solvent in initial experiments, but not all SOA were found to be fully soluble in water (Romonosky et al., 2015; Updyke et al., 2012). Therefore, methanol was used as the extracting solvent for the remaining experiments. To determine the extraction efficiency using methanol, filters were extracted twice in two separate aliquots of solvent, and absorption spectra were taken using a UV/Vis spectrometer, as described below. From comparison of the recorded absorbance values, it was concluded that extraction efficiency in the first extraction step was greater than 80% in all cases. We did not correct for the incompletely extracted material and assumed that the first step extracted everything; this assumption introduces up to 20% uncertainty in the resulting absorption coefficients. Typical concentrations of SOA in the extract were 20–900 μ g/mL, which translates into 70–3000 μ M, assuming an effective molecular weight of 300 g/mol for the SOA compounds.

The SOA extracts were placed in a 1 cm quartz cuvette for absorption measurements. A dual beam spectrometer (Shimadzu UV-2450), with pure solvent as the reference, was used to take the UV/Vis absorption spectra. Wavelength-dependent mass absorption coefficients (MAC), in units of cm² g⁻¹, were calculated from the base-10 absorbance, A₁₀, of an SOA extract with the solution mass

concentration, C_{mass} (g cm^{-3}), and path length, b (cm):

$$MAC(\lambda) = \frac{A_{10}^{solution}(\lambda) \times \ln(10)}{b \times C_{mass}} \quad (1)$$

Note that the mass-normalized absorption coefficients of the bulk SOA material calculated in Eq. (1) should not be confused with the mass-normalized absorption cross sections of aerosolized particles (Laskin et al., 2015). In order to calculate C_{mass} from the SMPS data we assumed an average density of 1.2 g/cm^3 for all the SOA samples we generated. A number of other values for the SOA densities have been reported (Hallquist et al., 2009). The following equation can be used to correct the MAC values reported in this work to the actual MAC values if a different SOA density is adopted:

$$MAC_{corrected}(\lambda) = \frac{MAC_{reported}(\lambda) \times 1.2 \text{ g/cm}^3}{\rho_{actual}} \quad (2)$$

where ρ_{actual} refers to any other assumed or determined density. If $\rho_{actual} > 1.2 \text{ g/cm}^3$, the resulting corrected MAC values would be smaller than the ones reported here, while $\rho_{actual} < 1.2 \text{ g/cm}^3$ would result in larger corrected MAC values. In either case we do not expect the adjustment to be larger than 20% based on the spread of the SOA densities listed in Hallquist et al. (2009). In combination with the 20% uncertainty for the completeness of SOA extraction, the overall accuracy of the reported MAC values should be on the order of $\pm 30\%$. This is sufficient for the purpose of estimation of lifetimes done in this paper.

For water-soluble SOA investigated in this study, the MAC values from the absorption spectra of methanol extracts and the water extracts were found to be comparable in magnitude, as seen in Fig. 1. For SOA that was poorly soluble in water, the shape of the absorption spectrum in water and in methanol were similar, but the methanol spectrum had higher absorbance because it contained more dissolved material. In view of the small sensitivity of the shape of the spectrum to the extracting solvent, it is expected that the absorption spectra of SOA compounds are not too sensitive to the environment, and MAC calculated from Eq. (1) approximate MAC of bulk SOA material reasonably well.

Eq. (1) is only valid in the low concentration limit in which the Beer–Lambert law is obeyed. The Supporting Information section contains examples of verification of linearity of the measured base-10 absorbance as a function of the mass concentration of SOA in the solution (Figs. S1–S3). Such absorbance linearity checks were carried out for all the SOA samples examined in this work, typically at 280 nm, where all SOA had easily detectable absorbance. Because absorbance values of the SOA solutions were recorded at different SOA concentration levels, the MAC spectra were built from the

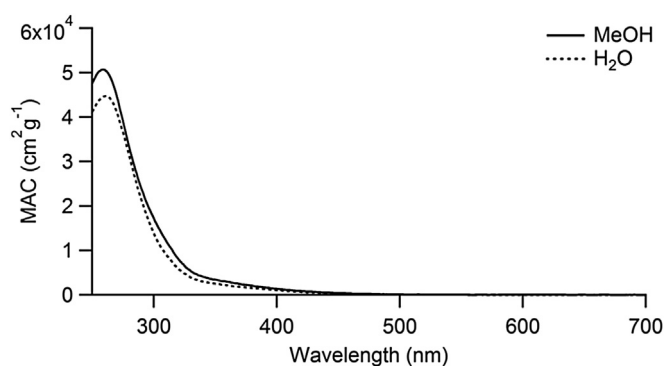


Fig. 1. A comparison of MAC values for 1MPYR/O₃ SOA, extracted in either water or methanol.

absorbance data in a piecewise fashion, wherein the solutions with higher concentrations were used to get MAC at longer wavelengths, and solutions with lower concentration were used to get MAC at shorter wavelengths. This helped improve the signal-to-noise ratio in MAC at longer wavelengths without having to violate the Beer–Lambert law at shorter wavelengths.

3. Results and discussion

The MAC values for all the SOA samples prepared in this study are shown on a logarithmic scale in Fig. 2. The same figure also shows the MAC values on a linear scale for the SOA derived from APIN, ISO, and XYL; the corresponding plots for the rest of SOA can be found in Figs. S4–S29 of the supporting information section. The MAC spectra have characteristic shapes for absorption spectra of common organic molecules, with the peak of absorption occurring in the UV range and a tail of absorption in the near-UV and visible ranges. The MAC values for some of the SOA samples, especially the ones generated from ISO, appear noisy because of the low SOA mass yield resulting in low measured absorbance values for the SOA solutions. In agreement with the observations of Liu et al. (2015), the MAC values for the SOA from aromatic precursors (TOL, XYL, GUA, NAP) are roughly an order of magnitude higher than MAC values for the SOA from isoprene and terpenes. Surprisingly, high MAC values were observed for SOA derived from heteroaromatic precursors 1MPYR (Fig. S22) and 2MPYR (Figs. S23–S25). It was initially suspected that the presence of the heterocyclic nitrogen atoms in these compounds may contribute to the high absorption coefficients of their SOA; however, ozonolysis of IMID resulted in SOA with relatively low MAC values (Fig. S12). Clearly, the absorption spectra of SOA produced from nitrogen containing organics are highly sensitive to the structure of the VOC precursor.

The rate constant of a photochemical reaction, J , for any compound in SOA can be calculated from

$$J = \int \sigma(\lambda)\phi(\lambda)F(\lambda)d\lambda \quad (3)$$

where σ is the absorption cross-section [$\text{cm}^2 \text{ molec}^{-1}$], ϕ is the quantum yield for the process of interest, and F is the spectral flux density [$\text{photon cm}^{-2} \text{ s}^{-1} \text{ nm}^{-1}$]. Whereas the gas-phase absorption cross-sections and quantum yields are well known for a number of simple organic compounds (Sander et al., 2011), there is only limited information available for the multifunctional molecules found in SOA. Furthermore, significant matrix effects on absorption spectra and quantum yields can be expected in the highly-viscous environment inside the particles (Lignell et al., 2014). Another complication is that a complex environmental mixture, such as an organic aerosol, contains a broad spectrum of compounds with different photophysical and photochemical processes. Some of them do not measurably absorb radiation in the wavelength of interest. Others absorb the radiation, but can dissipate the energy without undergoing a reaction, i.e., through internal conversion. Only a certain fraction of the components of aerosol can undergo photolysis, photoisomerization, photoassociation and other photochemical reactions leading to a change in the aerosol composition.

In this work, a simplified approach was taken in which an effective photochemical rate constant is calculated from sample-averaged properties. This approach assumes that the composition of the SOA can be represented by an average formula with an effective molecular weight (MW), and that all of the SOA molecules have identical photophysical properties. The effective MW used for this study was 300 g/mol, as determined by averaging MW from high-resolution mass spectra of broad range of SOA reported by

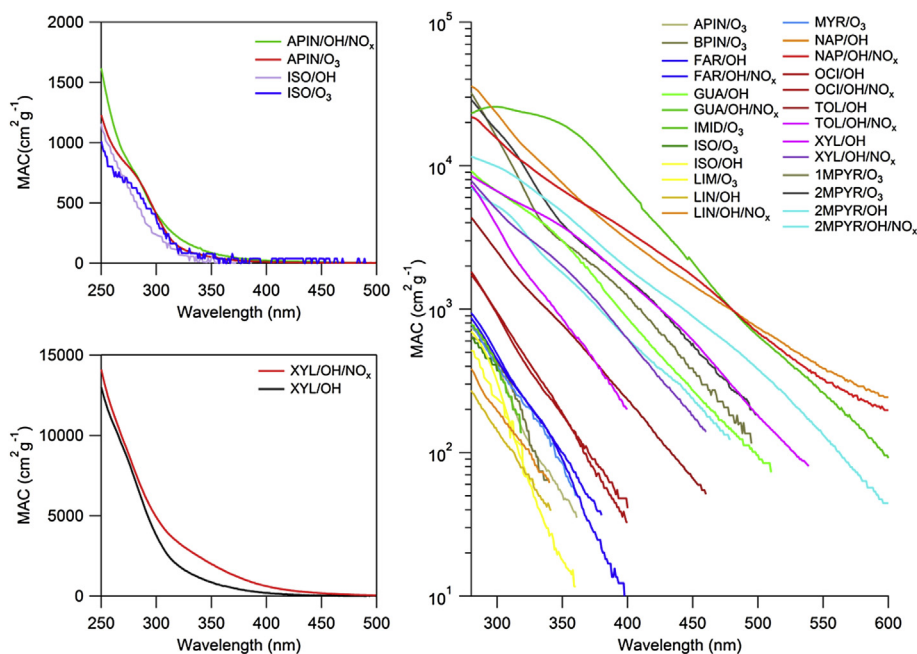


Fig. 2. The measured MAC values for all the SOA on a logarithmic scale. Two left panels compare MAC values for XYL, APIN, and ISO SOA on a linear scale.

Romonosky et al. (2015). The effective absorption cross section can then be estimated from MAC values (where N_A is the Avogadro's constant).

$$\sigma_{\text{effective}}(\lambda) = \frac{\text{MAC}(\lambda) \times \text{MW}}{N_A} \quad (4)$$

The quantum yields for the photochemical reactions of SOA compounds inside the particles are not known. As a limiting case, unity quantum yield is assumed, characteristic of the photolysis of hydrogen peroxide (Sander et al., 2011). Specifically, a step-function wavelength dependence of the quantum yield is used, in which $\phi = 1$ at all wavelengths below $\lambda_0 = 400$ nm, and $\phi = 0$ above λ_0 . The upper wavelength limit is introduced because the quantum yield can be expected to drop sharply below the dissociation threshold. For example, NO_2 photolysis occurs with $\phi \sim 1$ below 398 nm, and quantum yield declines rapidly above this threshold wavelength (Finlayson-Pitts and Pitts, 2000). Likewise, O_3 has a high quantum yield below its dissociation threshold at 310 nm, and it drops sharply above 310 nm.

The assumption of the unity value may be unrealistic for condensed-phase photochemical processes; therefore, a model in which the quantum yield decreases with increasing wavelengths is also considered. The wavelength dependence of the quantum yield of acetone at 1 atm is adopted as a more realistic case. Specifically, acetone undergoes efficient photolysis only below ~ 300 nm, where the solar spectral flux is quite small, and is resilient to photolysis at longer wavelengths. Similar behavior is expected from condensed-phase carbonyl compounds as well as other types of organic species. Fig. 3 shows the assumed quantum yields for both the H_2O_2 -like model and the acetone-like model.

With these assumptions, the effective photochemical rate constants can be calculated for the solar-zenith-angle (SZA) dependent values of solar spectral flux, F .

$$J_{\text{effective}} = \int \sigma_{\text{effective}}(\lambda) \phi(\lambda) F(\lambda, \text{SZA}) d\lambda \quad (5)$$

Fig. 4 shows the resulting dependence of J on SZA for 2MPYR/ O_3

(one of the more absorbing SOA in our data set) and ISO/OH (one of the less absorbing SOA). The peak values of J range from 10^{-2} to 10^{-4} s^{-1} (when using the H_2O_2 model for the quantum yields), which represent fast processes in comparison to typical atmospheric reactions. Fig. 5 shows the dependence of J on the time of day for Los Angeles (34° latitude) on June 22, 2014.

The effective lifetime for the photochemical change in SOA can be calculated as an inverse of the 24-h average rate

$$\tau = \left(\langle J_{\text{effective}} \rangle_{24\text{hr}} \right)^{-1} \quad (6)$$

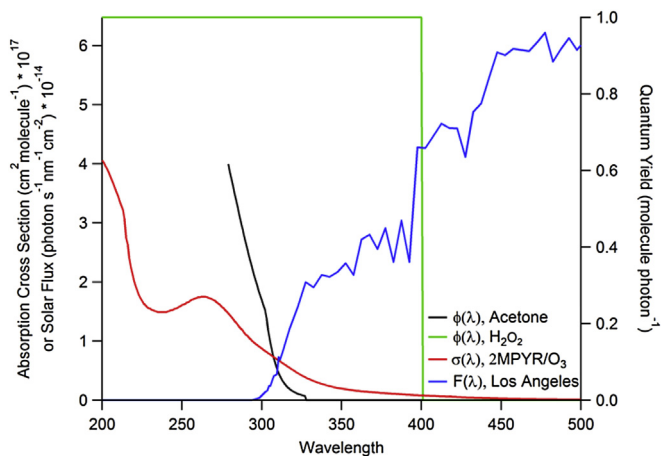


Fig. 3. The wavelength dependence of quantum yields used in this work: the H_2O_2 model with unity quantum yield below 400 nm (green trace) and the quantum yields for the photolysis of acetone at 1 atm (black trace). These quantum yields were convoluted with the effective absorption cross sections (an example for 2MPYR/ O_3 is shown in red trace) and solar spectral flux (an example at 13.00 h in Los Angeles in June is shown in blue trace) to obtain the effective photochemical rate constants. The absorption cross sections and the flux were multiplied by factors of 10^{17} and 10^{-14} , respectively, in order to show them in the same plot. (For interpretation of the references to color in this figure legend, the reader is referred to the web version of this article.)

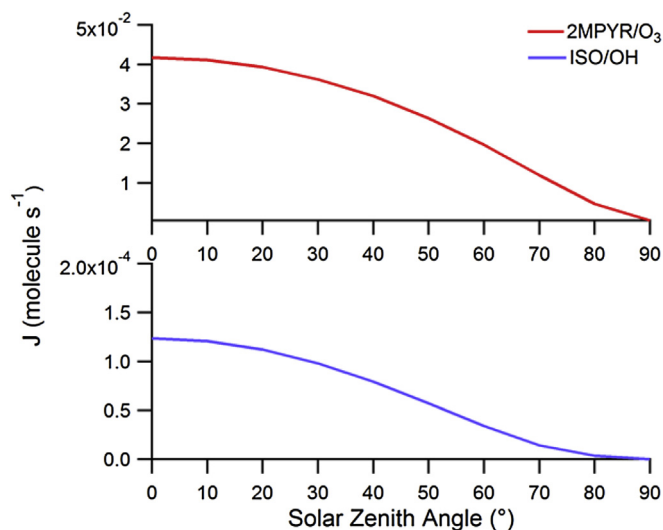


Fig. 4. Photochemical rate constant, J , of 2MPYR/O₃ SOA and ISO/OH SOA as a function of SZA calculated using Eq. (5) and assuming the H₂O₂ model for the quantum yields.

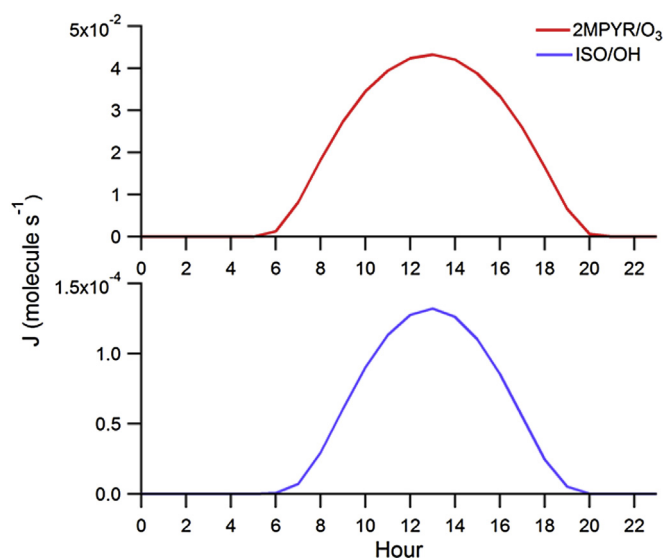


Fig. 5. Photochemical rate constant, J , for the same SOA samples as used in Fig. 4 plotted as a function of time over a 24 h period in Los Angeles (34° latitude) on June 22, 2014. The 24-h average values of $\langle J \rangle$ are listed in Table 3 for all the SOA types probed in this work.

The resulting values calculated from Eq. (6) for all the SOA types probed here for both the H₂O₂ and acetone models of quantum yields are summarized in Table 3. This table also lists the lifetimes during the peak irradiation in Los Angeles at 13:00 h, which is more meaningful for compounds with very short lifetimes.

The data in Table 3 reflect the high level of diversity in the optical properties of different types of biogenic and anthropogenic SOA. They also suggest that photochemical reactions in SOA can potentially occur on atmospherically meaningful time scales. Lifetime estimates for the SOA samples using the quantum yield of H₂O₂ are on the order of minutes, with the peak irradiation lifetimes being as short as a few seconds in some cases. These lifetimes are probably unrealistically short because it is unlikely that SOA compounds can photolyze with quantum yields approaching unity. The lifetime estimates relying on the quantum yield of acetone range from 60 to 11,000 min (1–180 h) and likely represent a more

realistic estimate for the time scales of photochemical reactions occurring in aerosols. For example, the lifetime predicted for APIN/O₃ SOA with the acetone model for quantum yields (4800 min or 3.3 days) is of the same order of magnitude as the lifetime for the photolysis-induced loss of peroxides from APIN/O₃ SOA (~10 days) measured by Epstein et al. (2014). It is noted that a direct comparison is complicated in this case by the possible photochemical recycling of peroxides in the particles in the Epstein et al. (2014) experiments. The predicted average lifetimes of 60–90 h for NAP SOA (with the acetone model for quantum yields) are also in fair agreement with the observed lifetimes for photobleaching of NAP SOA solutions in water, which ranged from ~15 to 90 h depending on the observation wavelength (Lee et al., 2014). Assuming that molecules in SOA can undergo photochemistry with photolysis quantum yields similar to that of acetone, condensed-phase photochemistry can become competitive with the more traditional mechanisms of SOA aging, such as multi-phase oxidation by OH. This is the most important conclusion of this work.

The SOA examined in this work was prepared under dry conditions and without using seed particles because we wanted to minimize contributions from the seed particle material to the measured absorption coefficients of the organic material in SOA. While neutral seed particles, such as ammonium sulfate, are not expected to affect the optical absorption properties of the SOA material, highly-acidic seed particles may affect both the mechanism of SOA production (Czoschke et al., 2003) and the optical absorption properties (Nozière and Esteve, 2007; Song et al., 2013; Van Wyngarden et al., 2015) of SOA. For example, Lin et al. (2014) found that SOA formed by uptake of isoprene epoxydiol (IEPOX) onto acidic particles was considerably more absorbing than that SOA formed on neutralized seed particles. To test for the effect of the seed particle acidity, we have included the IEPOX/acidic seed data from Lin et al. (2014) in Table 3. The estimated lifetime of the IEPOX/acidic seed SOA is more than an order of magnitude shorter than that for IOS/OH or ISO/O₃ SOA. Clearly, the effect of the aerosol seed acidity and the coupled effect of the relative humidity on the optical absorption properties of SOA are significant and deserve further study.

Tables 1 and 2 indicate that the amount of collected SOA material was typically between 0.1 and up to 3.5 mg. The large amount of SOA material was necessary for the accurate measurement of absorption coefficients. The high material requirements of these experiments necessitated using unrealistically high concentrations of VOC precursors, especially in the case of isoprene, and for all flow tube experiments. Future experiments should address the question whether the MAC values are strongly dependent on the VOC concentration used in the SOA preparation.

While our analysis predicts that condensed-phase photochemistry in SOA may be reasonably fast under typical atmospheric conditions, it does not say anything about the effect of these photochemical reactions on the climate- and health-related properties of aerosols, such as their mass concentration, average particle size, absorption coefficient, toxicity, etc. This effect will, of course, depend on the prevailing type of photochemistry occurring in SOA particles. For example, in the SOA derived from monoterpenes and isoprene, photolysis of peroxide and carbonyl functional groups can be expected to fragment SOA compounds into smaller molecules (Romonosky et al., 2015), which may then evaporate from particles to gaseous phase as observed in several previous experiments with bulk SOA samples (Hung et al., 2013; Mang et al., 2008; Walser et al., 2007). The evaporation will lead to SOA particle shrinking, as observed, for example in experiments by Epstein et al. (2014) and Wong et al. (2014) with APIN/O₃ SOA. However, it should not be assumed that fragmentation is always going to be the dominant photochemical pathway. For example,

Table 3
Effective SOA lifetimes with respect to photochemistry for Los Angeles.

Sample	Effective lifetime using QY of H ₂ O ₂ (24 h average) (min)	Peak effective lifetime using QY of H ₂ O ₂ (min)	Effective lifetime using QY of acetone (24 h average) (min)	Peak effective lifetime using QY of acetone (min)
1MPYR/O ₃	1.4	0.52	110	31
2MPYR/O ₃	1.0	0.39	85	24
2MPYR/OH	2.6	0.96	260	75
2MPYR/OH/NO _x	1.0	0.38	130	37
APIN/O ₃	85	30	4800	1300
MYR/O ₃	58	21	3800	1100
BPIN/O ₃	90	31	4000	1100
FAR/OH	53	19	3500	1000
FAR/OH/NO _x	47	17	3400	960
GUA/OH	1.7	0.63	190	56
IMID/O ₃	95	34	4800	1300
ISO/O ₃	88	32	5400	1500
ISO/OH	410	130	7600	2100
LIM/O ₃	230	77	7200	2000
LIN/OH/NO _x	100	37	7700	2200
LIN/OH	160	56	11,000	3100
OCI/OH	25	8.9	1700	480
OCI/OH/NO _x	25	9.1	1800	520
XYL/OH	5.5	2.0	430	120
XYL/OH/NO _x	2.5	0.9	280	81
Toluene/OH	6.2	2.3	590	170
Toluene/OH/NO _x	1.3	0.50	190	58
Naphthalene/OH	0.62	0.23	64	18
Naphthalene/OH/NO _x	0.68	0.25	89	26
Catechol/O ₃ ^a	3.0	1.1	260	75
IEPOX/MgSO ₄ seed particle/ dry conditions ^b	6.4	2.3	540	154

^a The MAC data were obtained from Liu et al. (2013) and used in lifetime calculations.

^b The MAC data were obtained from Lin et al. (2014) and used in lifetime calculations.

cis-pinonic acid, a major product of ozonolysis of APIN, undergoes a reasonably efficient photoisomerization into limononic acid, a structural isomer of cis-pinonic acid (Lignell et al., 2013). For the cases when photoisomerization is the dominant photochemical process, there should not be any effect on the particle size or mass concentration. Photoassociation reactions, in which two SOA compounds dimerize after absorbing a photon, will produce larger products than the initial reactants. This will have an effect of reducing the volatility of SOA compounds making the aerosol less likely to evaporate upon dilution. An even more interesting scenario is conceivable in which a photoexcited species on the particle surface reacts with a nearby VOC producing less volatile products and leads to an increase in the particle mass, as observed in recent laboratory experiments with aerosols containing photosensitizers (Aregahegn et al., 2013; Monge et al., 2012). In summary, condensed-phase photochemistry can either decrease or increase the particle mass and size depending on the nature of the dominant photochemical processes.

The effect of condensed-phase photochemistry on the optical properties of SOA is also very important because it has direct consequences for the radiative forcing by aerosols (Laskin et al., 2015). Light-absorbing SOA can lose or gain their chromophoric species as a result of various photochemical reactions. For example, recent laboratory experiments demonstrated that the wavelength dependence of the absorption coefficient can be dramatically altered by UV irradiation of several types of model brown carbon compounds (Lee et al., 2014; Rincon et al., 2010; Sareen et al., 2013; Zhao et al., 2015; Zhong and Jang, 2014). Such changes in the absorption coefficient of SOA are also important for gas-phase photochemistry because aerosols are capable of screening UV radiation and reducing ozone production in heavily polluted areas (Finlayson-Pitts and Pitts, 2000). Clearly, more work is needed to explore the rate and occurrence of such photochemical effects on the SOA absorption spectra before these effects can be incorporated in climate and air pollution models.

4. Conclusions

The optical properties of SOA representative of biogenic and anthropogenic atmospheric environments were examined by making model SOA in a chamber or a flow tube and measuring their bulk mass absorption coefficients (MAC). This information was used to estimate the effective absorption cross sections of SOA and lifetimes of SOA compounds with respect to possible condensed-phase photochemical reactions. The calculated lifetimes suggest that condensed-phase photochemistry may potentially compete with other known aerosol aging processes. These findings have important implications for understanding the aging of atmospheric aerosols, and the potential effect of this photochemical aging on climate- and health-related properties of SOA.

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Appendix A. Supplementary data

Supplementary data related to this article can be found at <http://dx.doi.org/10.1016/j.atmosenv.2015.10.019>.

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